N35 - SOLUTIONS

Colligative Properties

Ionic Solutes

Colligative properties

Properties whose value depends only on the number of solute particles, and not on what they are.

- **Value of the property depends on the concentration of the solution.**

The difference in the value of the property between the solution and the pure substance is generally related to the different attractive forces and solute particles occupying solvent molecules positions.

Colligative Properties

Colligative properties are those that depend on the concentration of particles in a solution, not upon the identity of those properties.

- **Boiling Point Elevation**
- **Freezing Point Depression**
- **Osmotic Pressure**

The van't Hoff Factor, *i*

Electrolytes may have two, three or more times the effect on boiling point, freezing point, and osmotic pressure, depending on its dissociation.

Ionic Comp.	$\mathbf{AB \rightarrow A^+ + B^-}$		$\mathbf{C_2D \rightarrow 2C^- + D^{2-}}$	
#particles	1	2	1	3
Nonelectrolyte	$\mathbf{XY(s) \rightarrow XY(aq)}$			
#particles	1	1		

Dissociation Equations and *i*

$i = 2$

$NaCl(s) \rightarrow Na^{+}(aq) + Cl^{-}(aq)$

$i = 2$

$AgNO_3(s) \rightarrow Ag^{+}(aq) + NO_3^{-}(aq)$

$i = 3$

$MgCl_2(s) \rightarrow Mg^{2+}(aq) + 2\ Cl^{-}(aq)$

$i = 3$

$Na_2SO_4(s) \rightarrow 2\ Na^{+}(aq) + SO_4^{2-}(aq)$

$AlCl_3(s) \rightarrow Al^{3+}(aq) + 3\ Cl^{-}(aq)$

$i = 4$

Ideal vs. Real van’t Hoff Factor

The ideal van’t Hoff Factor is only achieved in VERY DILUTE solution.

	Molality, *m*					
Solute	**1.0**	**0.10**	**0.010**	**0.0010**	**…**	**Inf dil***
NaCl	1.81	1.87	1.94	1.97	…	2
$MgSO_4$	1.09	1.21	1.53	1.82	…	2
$Pb(NO_3)_2$	1.31	2.13	2.63	2.89	…	3

Freezing Point Depression

Each mole of solute particles lowers the freezing point of 1 kilogram of water by 1.86 degrees Celsius.

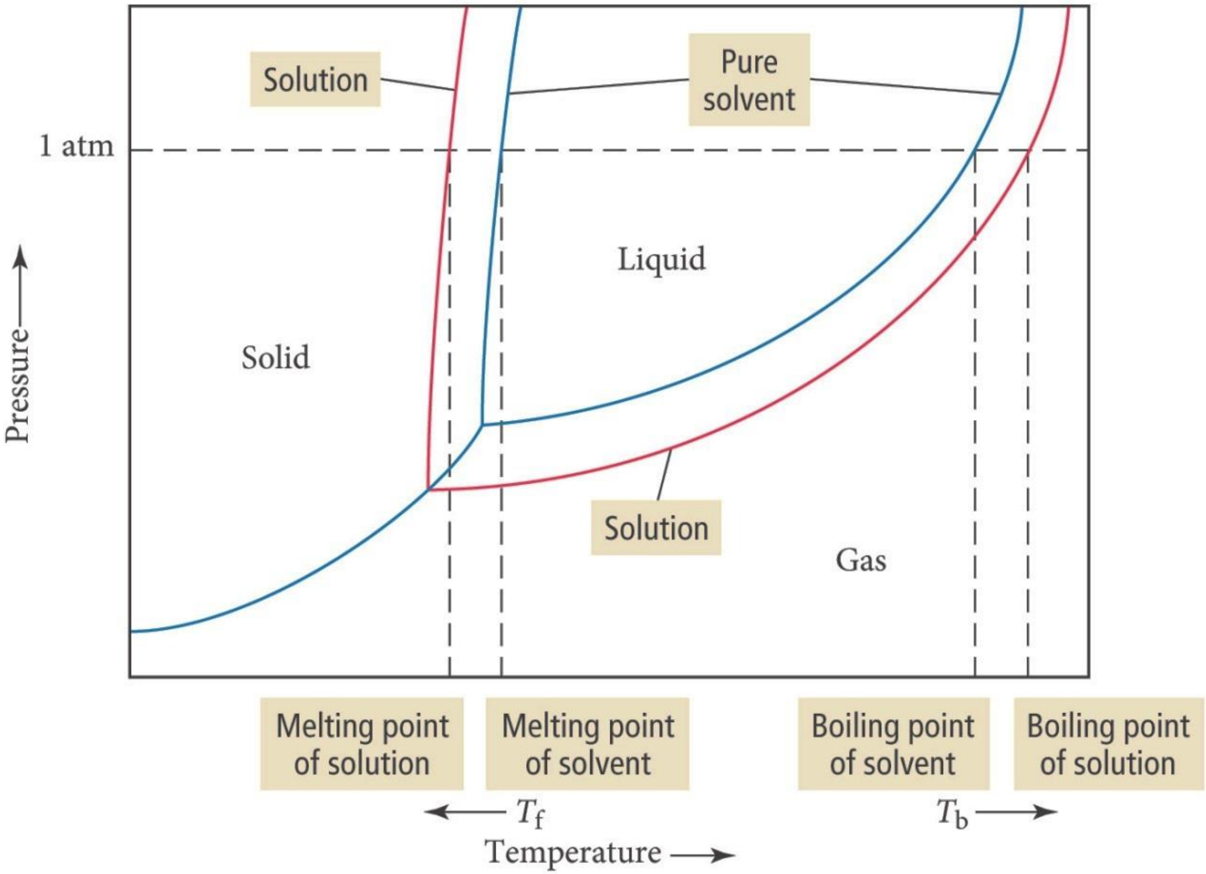
$$\Delta T = i \cdot K_f \cdot m_{solute}$$

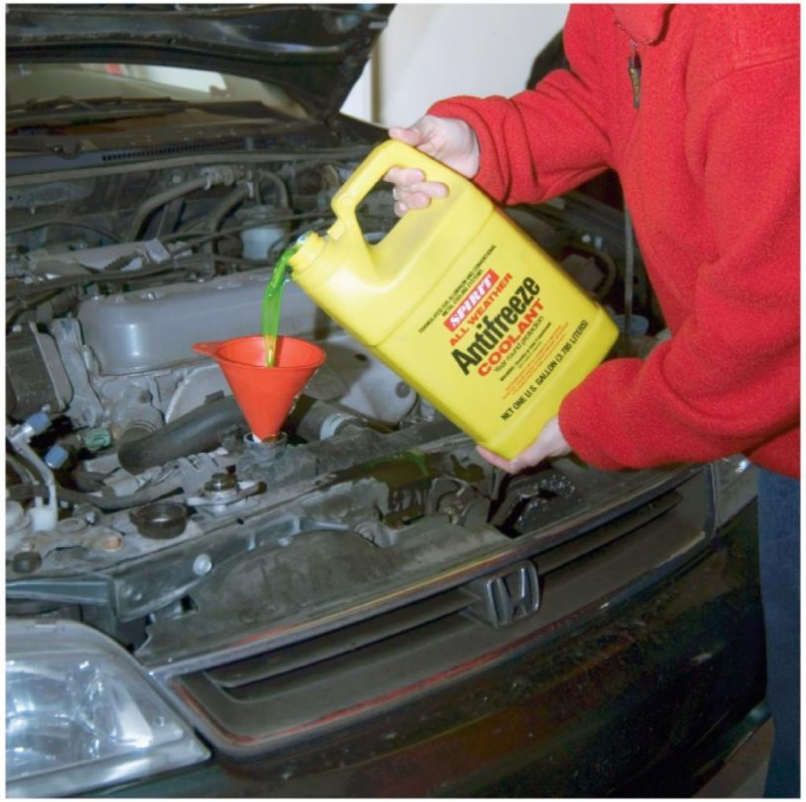
K_f **= 1.86 °C · kilogram/mol**

m **= molality of the solution**

i **= van’t Hoff factor**

Freezing Point Depression and Boiling Point Elevation





Boiling Point Elevation

Each mole of solute particles raises the boiling point of 1 kilogram of water by 0.51 degrees Celsius.

$$\Delta T = i \cdot K_b \cdot m_{solute}$$

K_f = 0.51 °C · kilogram/mol

m = molality of the solution

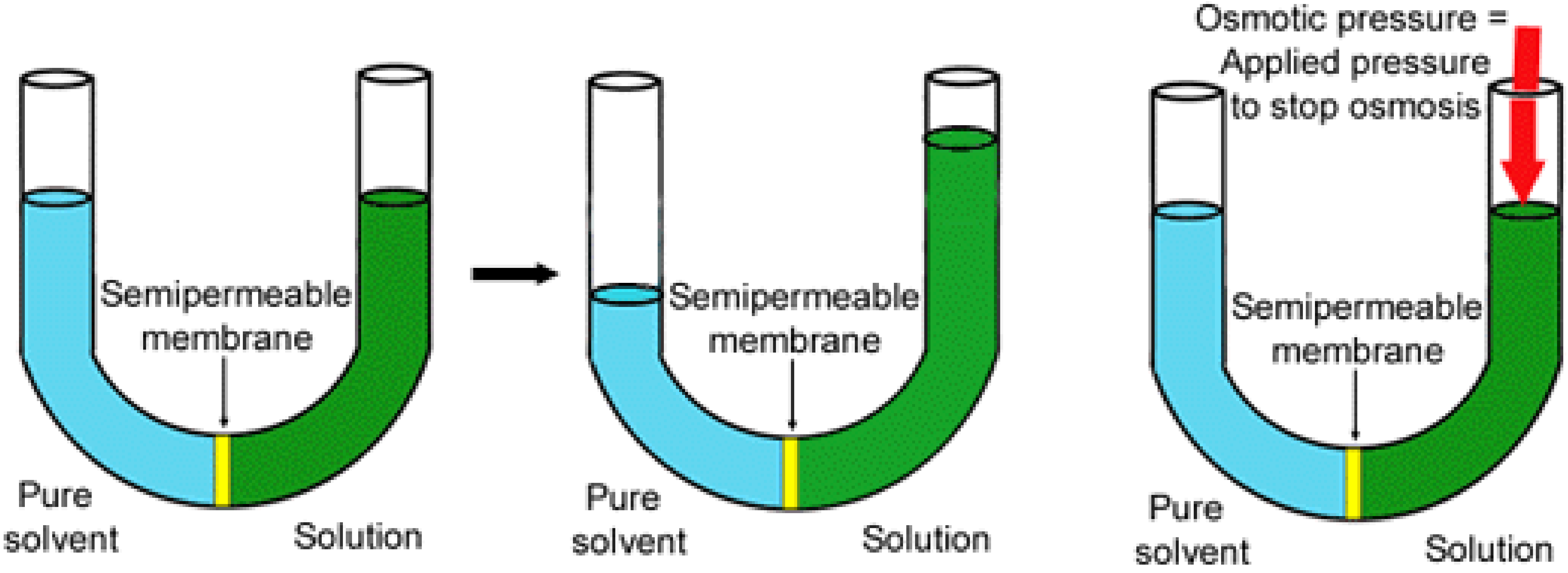
i = van’t Hoff factor

Freezing Point Depression and Boiling Point Elevation Constants, °C/m

Solvent	K_f	K_b
Acetic acid	3.90	3.07
Benzene	5.12	2.53
Nitrobenzene	8.1	5.24
Phenol	7.27	3.56
Water	1.86	0.512

Osmotic Pressure

The minimum pressure that stops the osmosis is equal to the osmotic pressure of the solution



Osmotic Pressure Calculations

$$\Pi = iMRT$$

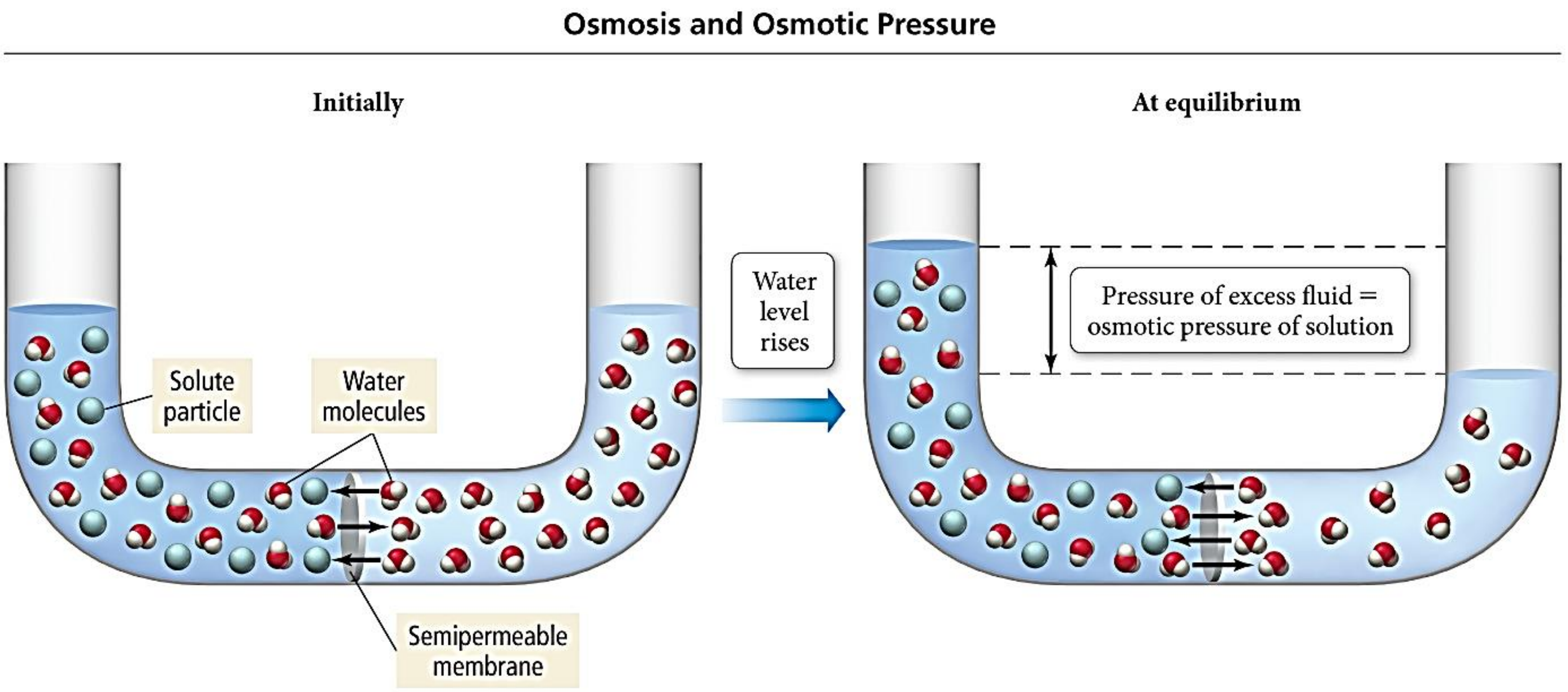
Π = Osmotic pressure

M = Molarity of the solution

R = Gas Constant = 0.08206 L·atm/mol·K

T = Absolute Temperature

An Osmosis Cell



Suspensions and Colloids

Suspensions and colloids are NOT solutions.

Suspensions - The particles are so large that they settle out of the solvent if not constantly stirred.

Colloids - The particles intermediate in size between those of a suspension and those of a solution.

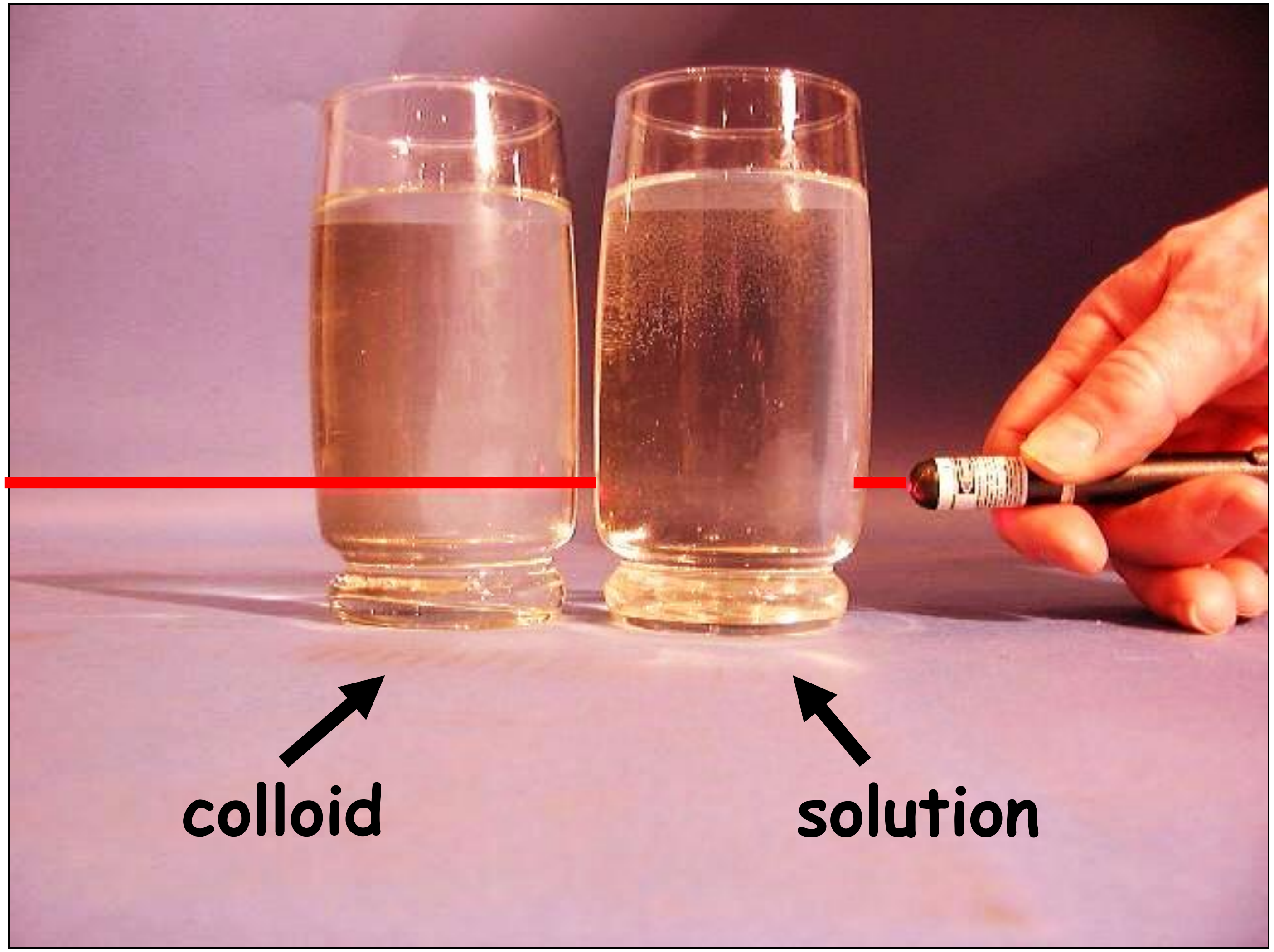
Types of Colloids

Examples	Dispersing Medium	Dispersed Substance	Colloid Type
Fog, aerosol sprays	Gas	Liquid	Aerosol
Smoke, airborn germs	Gas	Solid	Aerosol
Whipped cream, soap suds	Liquid	Gas	Foam
Milk, mayonnaise	Liquid	Liquid	Emulsion
Paint, clays, gelatin	Liquid	Solid	Sol
Marshmallow, Styrofoam	Solid	Gas	Solid Foam
Butter, cheese	Solid	Liquid	Solid Emulsion
Ruby glass	Solid	Solid	Solid sol

The Tyndall Effect

Colloids scatter light, making a beam visible. Solutions do not scatter light.

Which glass contains a colloid?



All of the following are colligative properties except:

- A Osmotic Pressure
- B Boiling Point elevation
- C Freezing Point depression
- D Density elevation
- E None of these

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A. Osmotic Pressure

B. Boiling Point elevation

C. Freezing Point depression

D. **Density elevation**

E. None of these

121.2 g of NaCl completely dissolve (producing Na^+ and Cl^- ions) in 1.00 Kg of water at 25°C. The VP of pure water at this temp is 23.8 torr. Determine the VP of the solution

- A 22.1 torr
- B 22.9 torr
- C 20.6 torr
- D 19.9 torr
- E 23.8 torr

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$$\frac{121.2\text{g} \;|\; 1\text{mol}}{\;|\; 58.44\text{g}} = 2.07 \text{ mol}$$

$$\frac{1000\text{g} \;|\; 1\text{mol}}{\;|\; 18.02\text{g}} = 55.5\text{mol}$$

$$X_{H2O} = \frac{55.5}{(2.07 + 55.5)} = 0.964$$

$P = (0.964)(23.8)$

= 22.9torr

At 40°C, heptane has a vapor pressure of about 92.0 torr and octane has a vapor pressure of about 31.2 torr. Assuming ideal behavior, what is the vapor pressure of a solution that contains twice as many moles of heptane as octane?

- **A** 61.3 torr
- **B** 51.5 torr
- **C** 71.7 torr
- **D** 82.1 torr
- **E** None of these

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Octane = 1 mol

Heptane = 2 mol

$P = X_1P_1 + X_2P_2$

$$P = \left(\frac{1}{3}\right)(31.2) + \left(\frac{2}{3}\right)(92.0)$$

P = 71.7 torr

A 7.12 g sample of a compound is dissolved in 250. grams of benzene. The freezing point of this solution is 1.02°C below that of pure benzene. What is the molar mass of this compound? ($K_{f\ benzene}$ = 5.12 °C/m). Ignore sig. figs.

- A. 35.7 g/mol
- B. 143 g/mol
- C. 286 g/mol
- D. 5.67 g/mol
- E. 71.5 g/mol

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- C 286 g/mol
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$\Delta T = iKm$

$1.02 = (1)(5.12)(m)$

$m = 0.199\ mol_{solute}/kg_{solvent}$

$$0.199 = \frac{x\ mol}{0.250\ kg} = 0.04975 \text{ moles}$$

mm = g/mol = 7.12g/0.04975mol

= 143g/mol

What is the boiling point change for a solution containing 0.681 moles of naphthalene (a nonvolatile, nonionizing compound) in 250. g of liquid benzene? (K_b = 2.53 °C/m for benzene)

A. 6.89 °C

B. 0.93 °C

C. 3.72 °C

D. 1.723 °C

E. 0.431 °C

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$\Delta T = iKm$

$m = mol_{solute}/kg_{solvent}$

$$= \frac{0.681\, mol}{0.250\, kg} = 2.724 \text{ m}$$

$\Delta T = (1)(2.53)(2.724) = 6.89°C$

Determine the Osmotic Pressure of a solution that contains 0.017 g of a hydrocarbon solute (molar mass = 340 g/mol) dissolved in benzene to make 350-ml solution. The temperature is 20.0°C

A. 0.18 torr

B. 0.9 torr

C. 1.2 torr

D. 2.6 torr

E. 2.4 torr

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$$\Pi = iMRT$$

0.017g	1mol
	340g

$= 5x10^{-5}$ mol

$(5x10^{-5}mol)/(0.350L) = 1.43x10^{-4}$ M

$$\Pi = (1)(1.43x10^{-4})(62.4)(293)$$

= 2.6 torr

Can stop here!

Henry's Law not covered anymore. You can keep going if interested though!

Henry’s Law (removed from AP)

- The solubility of a gas (S_{gas}) is directly proportional to its partial pressure, (P_{gas}).

$$S_{gas} = k_H P_{gas}$$

- k_H is called the **Henry’s law constant**.

TABLE 12.4 Henry’s Law Constants for Several Gases in Water at 25 °C

Gas	k_H (M/atm)
O_2	1.3×10^{-3}
N_2	6.1×10^{-4}
CO_2	3.4×10^{-2}
NH_3	5.8×10^{1}
He	3.7×10^{4}

A correct statement of Henry's law is:

A. the concentration of a gas in solution is inversely proportional to temperature.

B. the concentration of a gas in solution is directly proportional to the mole fraction of solvent.

C. the concentration of a gas in solution is independent of pressure.

D. the concentration of a gas in a solution is inversely proportional to pressure.

E. none of these

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The solubility of O_2 in water is 0.590g/L at an oxygen pressure of around 14.7 atm. What is the Henry's Law constant for O_2 (in units of mol/L·atm)?

- A) $4.01E^{-2}$
- B) $1.25E^{-3}$
- C) $7.97E^{2}$
- D) $2.71E^{-1}$
- E) None of them are within 5% of the correct answer

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- E None of them are within 5% of the correct answer

S = kP

0.590g	1mol
	32g

=0.0184 mol

(0.0184 mol/L) = K (14.7atm)

K = 1.25E-3